St. Xavier's School

Chemistry. Class X

PERIODIC TABLE

Date : 18/5/2020

Q. Fill in the blanks in the following statements, with suitable words :

Chapter 1. The Periodic Properties and their Variations

1. The modern Periodic Table has ..... periods.
2. Each ..... in the Periodic Table is comprised of elements having the same number of electron shells.
3. Elements in a period, all have the same number of ..... in their atoms.
4. Elements in a group, all have the same number of ..... .
5. The most active metals are located in ..... 1 and 2 of the Periodic Table.
6. The most reactive non-metals comprise group ..... of the Periodic Table.
7. The elements of ..... are known as typical elements.
8. The elements occupying left and right side groups of Periodic Table are called ..... elements.
9. The rare gases are placed in ..... group at the end.
10. Elements from atomic number 57 to ..... occupy same place in Periodic Table. These elements are called .....
11. Actinides are the elements from atomic number ..... to ..... and are radioactive.
12. The actinides and ..... are kept outside the Periodic Table to mark their peculiar properties.
13. The properties of elements are periodic function of their.....
14. The atomic size ..... as we move left to right across the period, because the ..... increases but the .... remains the same.
15. The metallic character ..... in a group as one moves from top to bottom.
16. The metallic character ..... in a period as one moves from right to left.
17. In a period or in a group, the larger the atomic size of an element, the..... metallic is the element.
18. Moving across a ..... of the periodic table, the elements show increasing ..... character.
19. The amount of energy involved in the reaction $X + \text{energy} \rightarrow X^+ + C^-$ is known as the .....of the element X.
20. Across a period, the ionization potential.....
21. Down the group, electron affinity.....
22. The higher the electron affinity of a non-metal, ... the chemically reactive the non-metal is.
23. The tendency to gain an electron ..... on moving down a group and ..... on moving across a period in the Periodic Table.
24. Elements having high ionization potential have ..... electron affinities.
25. The electronegativity of elements ..... across a period and ..... down a group.
26. In general, non-metals are ..... electronegative than metals.
27. On moving from left to right in a given period, the number of shells ..... .
28. Element X belongs to group 2 and period 3 of the Periodic Table. It has ..... electrons in the outer most shell.
29. Each period except period 1 in the Periodic Table begins with ..... and ends up with a ..... .
30. The metallic and non-metallic character depends upon the.....and..... of the elements.

Ans. 1. 7
2. period
3. electron shells
4. outer electrons
5. groups
6. 17
7. period 3
8. representative
9. zero
10. 71, lanthanides
11. 90, 103
12. lanthanides
13. atomic number
14. Decrease, nuclear charge, number of shells
15. Increases
16. Decreases
17. more
18. period, non-metallic
19. ionization potential
20. increases
21. decreases
22. more
23. increases; decreases
24. high
25. increases; decreases
26. more
27. remains the same
28. 2 (two)
29. an alkali metal; noble gas
30. atomic size, ionization potential